

## Properties Of Gases Section Review Answer

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### Properties Of Gases Section Review

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### 12.1 section Review - The properties of gases & 12.2 ...

Gases are easily 1 because of the 2 between particles in a' gas. The four variables used to describe a gas are pressure, (P), 3 4 2 S 5. 8. dec a g h5 No are c n and number of (n). You can use 6 theory to predict and explain how gases will respond to a change in conditions. Doubling the amount of gas in a rigid container 7 the pressure.

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### Properties Of Gases Section Review Answer Key

The next portion of this section describes the properties of gases (number of gas particles, volume, temperature, and gas pressure) with an emphasis on gas pressure. You will discover what gas pressure is and what causes it. The rest of this section shows how the ideal gas model can be used to explain the relationships between the properties.

### Chapter 11 - Gases

CHAPTER 11 REVIEW Gases SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. Pressure surf f a o c r e ce area. For a constant force, when the surface area is tripled the pressure is (a) doubled. (b) a third as much. ... and Their Properties Ideal Gases Properties of Gases ...

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Pressure. Amount in Moles. Ideal Gases. considered to be a "point mass". Has elastic collisions ( do not attract or repel each other) Are constant, random straight-line motions. Have an average Kinetic Energy directly related to temperature. Point Mass. a particle so small, its mass is nearly zero.

### Chapter 14: The Properties of Gases Flashcards | Quizlet

- These are often referred to as vapors.
- Many of the properties of gases differ from those of solids and liquids:
- Gases are highly compressible and occupy the full volume of their containers.
- When a gas is subjected to pressure, its volume decreases.
- Gases always form homogeneous mixtures with other gases.

### Chapter Ten- Gases #2 Pg 432 #5, 43, 45, 47, #3 Pg 432 #6 ...

between them are overcome, gases do not have a definite shape or a definite volume. The movement of particles and the collisions between them cause gases to diffuse. Diffusion is the spreading of particles throughout a given volume until they are uniformly distributed. Diffusion occurs in solids and liquids but occurs most rapidly in gases.

### Chapter 16: Solids, Liquids, and Gases

Properties of Gases The ideal gas model is used to predict changes in four related gas properties: volume, number of particles, temperature, and pressure. Volumes of gases are usually described in liters, L, or cubic meters, m<sup>3</sup>, and numbers of particles are usually described in moles, mol.

### Chapter 11 Gases

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at a pressure of 55 kPa and a 'gas law review sheet north bergen school district june 13th, 2018 - gas law review sheet remember to have the same units before you multiply your final Gas Laws Review Answers - stage-hotel.travelshop.vn Identify the properties of an ideal gas vs. a real gas .

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that are introduced in this section. Each blank can be completed with a term, short phrase, or number. Gases are easily , or squeezed into a smaller volume 1. because of the between particles in a gas. The four variables 2. used to describe a gas are pressure, (P), (V), (T), 3. and number of (n). 4. You can use theory to predict and explain how gases 5.

### 05 CTR ch14 7/12/04 8:13 AM Page 347 THE PROPERTIES OF ...

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